**Solutions and Equilibrium Remediation**

**Molarity and dilution questions**

Solve the following molarity/dilution problems. Show all work and include correct units!

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| 1. Find the molarity of a solution in which .58 moles of NaCl are dissolved in 2.5 L of solution.
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| 1. What is the original molarity of .50 L of a KMnO4 solution that is diluted to 2.00 L of a 0.500*M* solution?
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| 1. What volume of 0.25*M* solution can be made from .125 mol of KCl?
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| 1. Find the molarity of a 450 mL solution containing 13.7 g of ZnSO4.
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| 1. How many L of a .25 M CuCl2 solution can be made from 0.30 L of a 6 M CuCl2 solution?
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**Use the solubility Curve to answer the following questions:**

1. a. At 90°C, you dissolved 10 g of KCl in 100. g of water. Is this solution saturated or unsaturated?

 b. How do you know?

2. A mass of 100 g of NaNO3 is dissolved in 100 g of water at 80ºC.

a) Is the solution saturated or unsaturated?

3. What mass of solute will dissolve in **100 g** of water at the following temperatures?

* 1. KNO3 at 70°C \_\_\_\_\_\_\_\_\_\_\_\_
	2. NaCl at 100°C \_\_\_\_\_\_\_\_\_\_\_\_
	3. NH4Cl at 90°C \_\_\_\_\_\_\_\_\_\_\_\_

Equilibrium Practice:

Write an equilibrium expression for each of the following reactions.

1) 3O2 ⇌ 2O3

What happens to the equilibrium of this reaction if pressure is increased?

2) N2 + 3H2 ⇌ 2NH3

What happens to the equilibrium of this reaction if more nitrogen is added?

3) H2 + I2 ⇌ 2HI

What happens to the equilibrium of this reaction if Iodine is taken away?

4) PCl5 ⇌ PCl3 + Cl2