**Stoichiometry Remediation**

1. How many grams are in 2.30 mol of SO2?
2. How many moles are in 575 g of KClO3?
3. What are the empirical and molecular formulas for a compound with 83.625% carbon and 16.375% hydrogen and a molecular weight of 388.78 g/mol?
4. Calculate the molar mass for the following compounds:
	1. Al2(SO4)3

* 1. FePO4
1. Calculate the percent composition of each element in H2SO4.
2. How much BaSO4 can be formed from 196.0 g of H2SO4?

 BaO + H2SO4 ---> BaSO4 + H2O

1. Oxygen gas is sometimes prepared in labs by the thermal decomposition of potassium chlorate (KClO3). The balanced chemical equation is as follows: 2 KClO3(s) ---> 2 KCl(s) + 3 O2(g)

If 5.150 grams of KClO3 decompose, what volume of O2 would be obtained at STP?

1. a. How many molecules are in 3.45 mol of H2O2?

b. How many atoms of Oxygen are there?